

Periodic Table

1A		2A												8A					
1	H	4	Be											2	He				
1.00794		9.01218												4.00260					
3	Li	6	C											9	F				
6.941		12.011												18.998403					
11	Na	12	Mg											17	Cl				
22.98977		24.305												35.453					
19	K	20	Ca											16	S				
39.0983		40.08												32.066					
37	Rb	38	Sr											34	Se				
85.4678		87.62												78.96					
55	Cs	56	Ba											52	Te				
132.905		137.33												127.60					
87	Fr	88	Ra											84	Po				
(223)		226.025												(209)					
				3B		4B		5B		6B		7B		8B		1B		2B	
				21	22	23	24	25	26	27	28	29	30						
				Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn						
				44.9559	47.88	50.9415	51.996	54.9380	55.847	58.9332	58.69	63.546	65.39						
				39	40	41	42	43	44	45	46	47	48						
				Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd						
				88.9059	91.224	92.9064	95.94	(98)	101.07	102.906	106.42	107.868	112.41						
				71	72	73	74	75	76	77	78	79	80						
				Lu	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg						
				174.967	178.49	180.948	183.85	186.207	190.2	192.22	195.08	196.967	200.59						
				103	104	105	106	107	108	109									
				Lr	Unq	Unp	Sg*	Uns	Uno	Une									
				(260)	(261)	(262)	(263)	(262)	(265)	(266)									

Atomic number — 6
 Symbol — C
 Atomic mass — 12.011

57	La	58	Ce	59	Pr	60	Nd	61	Pm	62	Sm	63	Eu	64	Gd	65	Tb	66	Dy	67	Ho	68	Er	69	Tm	70	Yb
138.906	140.12	140.908	144.24	(145)	150.36	151.96	157.25	158.925	162.50	164.930	167.26	168.934	173.04	173.04	173.04	173.04	173.04	173.04	173.04	173.04	173.04	173.04	173.04	173.04	173.04	173.04	173.04
89	Ac	90	Th	91	Pa	92	U	93	Np	94	Pu	95	Am	96	Cm	97	Bk	98	Cf	99	Es	100	Fm	101	Md	102	No
227.028	232.038	231.036	238.029	237.048	(244)	(243)	(247)	(251)	(252)	(257)	(258)	(259)	(259)	(259)	(259)	(259)	(259)	(259)	(259)	(259)	(259)	(259)	(259)	(259)	(259)	(259)	(259)

* The name of element 106 has not yet been certified.

Atomic Number, Mass Number, Isotopes

1. Give the information...



element: _____

protons: _____

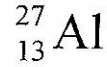
neutrons: _____



element: _____

atomic #: _____

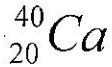
mass #: _____



element: _____

protons: _____

electrons: _____

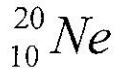


element: _____

protons: _____

average # neutrons: _____

atomic weight: _____



element: _____

atomic #: _____

electrons: _____

atomic weight: _____



element: _____

protons: _____

electrons: _____

avg. # neutrons: _____



element: _____

atomic #: _____

mass #: _____

electrons: _____



element: _____

atomic #: _____

electrons: _____

neutrons: _____



element: _____

protons: _____

electrons: _____

neutrons: _____

Atomic Structure Worksheet

Fill in the blanks for the elements in this chart. For the purposes of this chart, round all atomic masses to the nearest whole number.

Element	Number of Protons	Number of Neutrons	Number of Electrons	Atomic Mass	Atomic Number
lithium					
carbon					
chlorine					
silver					
lead					
calcium					
tantalum					
radium					
samarium					
uranium					
americium					
lawrencium					

The Atoms Family
Atomic Math Challenge

Name _____

<p>8 O Oxygen 15.999</p>	<p>← _____</p> <p>← _____</p> <p>← _____</p> <p>← _____</p>
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Atomic number equals
the number of

or

Atomic mass equals
the number of

+

<p>8 O _____ 15.999</p>

Atomic # = _____

Atomic Mass = _____

of Protons = _____

of Neutrons = _____

of Electrons = _____

<p>30 _____ Zinc 65.39</p>
--

Atomic # = _____

Atomic Mass = _____

of Protons = _____

of Neutrons = _____

of Electrons = _____

<p>3 Li _____ 6.941</p>

Atomic # = _____

Atomic Mass = _____

of Protons = _____

of Neutrons = _____

of Electrons = _____

<p>14 _____ Silicon 28.086</p>
--

Atomic # = _____

Atomic Mass = _____

of Protons = _____

of Neutrons = _____

of Electrons = _____

<p>5 B _____ 10.81</p>

Atomic # = _____

Atomic Mass = _____

of Protons = _____

of Neutrons = _____

of Electrons = _____

<p>35 _____ Bromine 79.904</p>
--

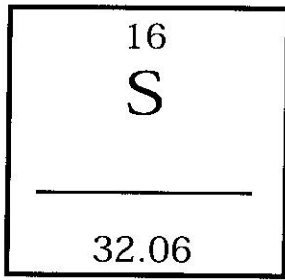
Atomic # = _____

Atomic Mass = _____

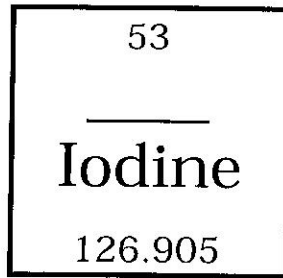
of Protons = _____

of Neutrons = _____

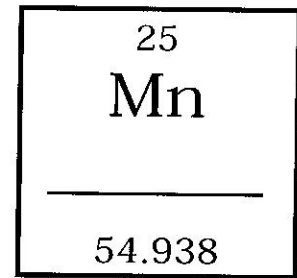
of Electrons = _____



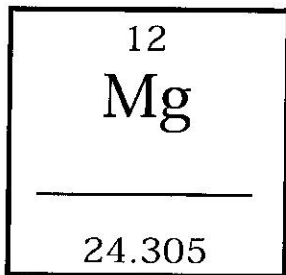
Atomic # = _____
 Atomic Mass = _____
 # of Protons = _____
 # of Neutrons = _____
 # of Electrons = _____



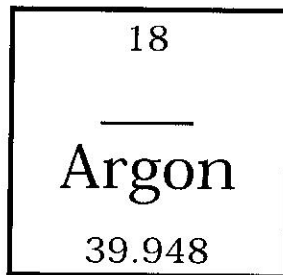
Atomic # = _____
 Atomic Mass = _____
 # of Protons = _____
 # of Neutrons = _____
 # of Electrons = _____



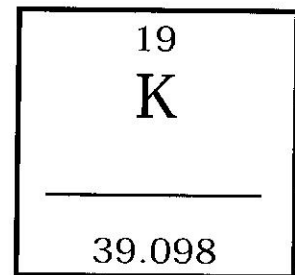
Atomic # = _____
 Atomic Mass = _____
 # of Protons = _____
 # of Neutrons = _____
 # of Electrons = _____



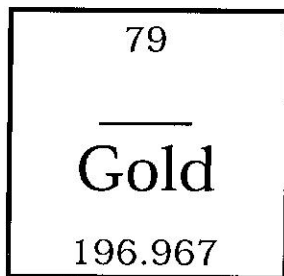
Atomic # = _____
 Atomic Mass = _____
 # of Protons = _____
 # of Neutrons = _____
 # of Electrons = _____



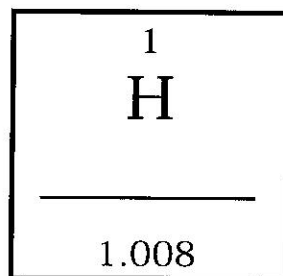
Atomic # = _____
 Atomic Mass = _____
 # of Protons = _____
 # of Neutrons = _____
 # of Electrons = _____



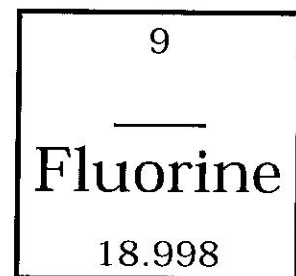
Atomic # = _____
 Atomic Mass = _____
 # of Protons = _____
 # of Neutrons = _____
 # of Electrons = _____



Atomic # = _____
 Atomic Mass = _____
 # of Protons = _____
 # of Neutrons = _____
 # of Electrons = _____



Atomic # = _____
 Atomic Mass = _____
 # of Protons = _____
 # of Neutrons = _____
 # of Electrons = _____



Atomic # = _____
 Atomic Mass = _____
 # of Protons = _____
 # of Neutrons = _____
 # of Electrons = _____

Protons, Neutrons, and Electrons Practice Worksheet

Fill in the blanks in the following worksheet. Please keep in mind that the isotope represented by each space may NOT be the most common isotope or the one closest in atomic mass to the value on the periodic table.

Atomic symbol	Atomic number	Protons	Neutrons	Electrons	Atomic mass
B			6		
	11				24
		31	37		
				39	89
	29		35		
		43			100
Pb					207
			102	70	
		89			225
Mo			53		
	81				206
	100		159		
No					261
Yb					172
		106	159		